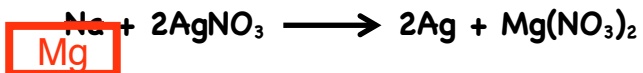


Single Displacement and Double Displacement Reactions

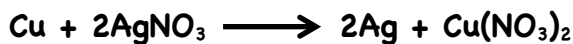
1. magnesium + silver nitrate \longrightarrow silver + magnesium nitrate



2. bromine + calcium iodide \longrightarrow iodine + calcium bromide



3. copper + silver nitrate \longrightarrow silver + copper (II) nitrate



4. zinc + hydrogen chloride \longrightarrow hydrogen + zinc chloride



5. lead (II) nitrate + potassium iodide \longrightarrow lead (II) iodide + potassium nitrate



6. calcium carbonate + hydrogen chloride \longrightarrow hydrogen carbonate + calcium chloride



7. aluminum + copper (II) chloride \longrightarrow copper + aluminum chloride



8. beryllium fluoride + sodium oxide \longrightarrow beryllium oxide + sodium fluoride



9. For each of the following, predict what the products are and then balance the equation.

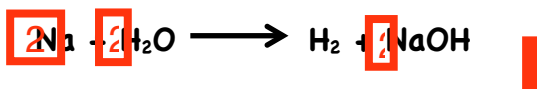
a) aluminum reacts with iron (III) oxide

(This is how iron is extracted from iron ore. It is heated at a high temperature with aluminum present. In the reaction, iron is separated from the iron ore producing a new aluminum compound.)



b) sodium reacts with water

(This reaction is dangerous. Dropping a piece of sodium in water can cause a fire.)



c) sodium chloride reacts with hydrogen sulfate (sulfuric acid)

(When this happens in the laboratory, a dangerous gas is produced.)



d) hydrogen sulfate reacts with calcium phosphate

One of the products is hydrogen phosphate. This is a very good plant fertilizer.)

